


How to calculate half life of iodine 131

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Half life questions

- The half-life of strontium-90 is 27 years. The half-life of sodium-24 is 15 hours.
 - 1. Sketch two curves on one set of axes to show how the number of atoms of each change with time.
- Two samples are prepared, one containing 10^{20} atoms of strontium and the other containing 10^{20} atoms of sodium.
- 2. Which of the two samples has the highest activity?
 - 3. A sample of iodine-131, with half-life 8.0 days, has an activity of 7.4×10^7 Becquerel (Bq). Calculate the activity of the sample after 4 weeks and 4days?
 - 4. ^{234}Th has a half life of 24.1 days.
 - (a) What fraction of a sample remains after 96.4 days?
 - (b) What fraction of a sample remains after 241 days?

$$A = \lambda N$$
$$N = N_0 e^{-\lambda t}$$
$$A = \lambda N_0 e^{-\lambda t}$$
$$N = \frac{A}{\lambda} e^{\lambda t}$$
$$N = \frac{A}{\lambda} e^{\lambda t}$$



Why is the half life of iodine 131 important. What is iodine 131 half life. How long does iodine 131 take to decay. How to find the half life of iodine 131.

Iodine 131 is a radioisotope with a very short half-life of 8.02 days, which makes it highly radioactive. Frequently used in small doses in thyroid cancer therapies, it is also one of the most feared fission products when accidentally released into the environment. Radiotoxicity of iodine 131 Radiotoxicity of iodine 131 is measured by a "interference dose converting factor" that allows the effective dose resulting from the ingestion of a given activity of a radio element to be calculated. The intake of Iodo131 is dangerous because it mainly affects the thyroid gland that plays a key role in child development. The toxicity of radioactive iodine varies greatly with age, with babies and small children and adolescents much more sensitive than adults. IN2P3 In medicine, iodine 131 is used primarily to study the operation of the thyroid, although also it may be used to treat hyperthyroidism as well as thyroid cancer. The first iodine production 131 in France took place in 1949 at Fort Chatillon, the site of the first atomic reactor ZOE, before manufacture was transferred to the Saclay nuclear research centre. The isotope was used since 1942, however, in the treatment of thyroid cancer. Although used in low doses for hand exams, iodine 131 is an ideal route for use in humans. Only a few radioactive atoms need to be inserted into the bloodstream so that the iodine path can be accurately controlled. The reactions that eventually transform into thyroid hormones; This is particularly interesting, given that iodine binds exclusively to the thyroid gland. Gamma-ray scintigraphy can thus monitor thyroid activity and signal the appearance of any anomalies. Over the past few years, iodine 131 has been abandoned in favour of another top, iodine 132-a gamma emitter with a half-life of only 13.2 Strongest doses of iodine 131 are also used in radioactive therapies aimed at dealing with Dogs. The iodine is inserted into the bloodstream in the same way, and the short trajectory of the emitted beta parts ensures that the radiation affects only a relatively small part of the body. Iodine 131 is also a dreaded fission product, posing as the main risk of short-term contamination in case of accidental release of waste. From a chemical point of view, iodine is a halogen (similar in structure to chlorine and fluorine) and its high volatility means that it easily turns into a purple vapor. Spinach contaminated by iodine-131spinach grown in northeastern Japan were contaminated by radioactivity from the Fukushima accident in March 2011. Radioactive dust deposits iodine-131 atoms in broad spinach leaves, leading to its sale ban. The iodine-131 is the most feared radioactive release after a nuclear accident due to its fixation by the soil. But its counterpart of its high activity - the amount of iodine-131 is divided by 2 every eight days, by 2500 every three months. After one month, no trace of iodine remains in food. Eugene Hoshiko / AP Volatile and highly mobile in the environment as volatile, radioactive iodine isotopes follow the usual processes of transfer once to the food chain: dispersion, deposition, absorption by plant leaves, absorption of the root, ingestion by animals and humans. Ingested by animals during lactation, the iodine deposited in the grass finds its way into milk a few hours after ingestion, the maximum that appears after three days. After a nuclear or radioactive accident, iodine-131 should be monitored in the food chain in previous weeks, until radionuclide disappears, especially in milk and vegetables, especially large vegetable leaves like spinach and lettuce. The water should also be monitored. In any case, the high radioactivity of iodine 131 is somewhat compensated by its high decay rate, with the level of activity falling by a factor of 1000 to eighty days. There are also procedures to protect ourselves from before it rots. Other radioactive iodine isotopes have very short lives, such as iodine-132 and iodine-133, whose periods are 20.8 and 2.3 hours. These isotopes deliver almost all their radioactivity in the first days after the closure of a reactor. Radioactive decomposition of short-lived iodine radioisotopes, in addition to radioisotopes of iodine-131, two other radioactive isotopes of iodine-131 and 132 (20.8 hours and 2.3 hours of duration). Iodine-132 comes from the decay of the tellurium-132 (period 3.2), using preliminary data published on the activities released after the Fukushima accident, the value displays the activity of the iodine radioisotopes during the first seven weeks in a row. The releases in the environment occurred from one to three days after the tsunami and the reactors were terminated. Iodine-133 had almost disappeared at the time of the first releases. The iodine-132 that closely follows the decomposition of the tellurium disappears within three weeks. At the end of these seven weeks, only remains iodine-131 which decreased 38 times. IN2P3 Iodine 129, another isotope of iodine, is one of the long-lasting fission products that have to be taken into account when dealing with radioactive waste. The semi-life of iodine 129 is 15.7 million years. Access to the French page

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