Alkyl halides and aryl halides class 12 notes

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haloalkanesalkyl halides: the general formula is rx, where r = group alchil. classification: general methods of prepared by alcohol in various ways: - action of hydrogen halogenurs: > the reactivity of halogenurs: > the re :allyl benzyl > tertiary > primary > alogenur of phosphorus : - action of thionium chloride (darzen process) This is known as Finkelstein reaction. - Fluoalkanes can also be prepared by alogenur action on ethers : >> Adding Hydrogen Alchere: Alkenes add on a hydrogen halogenur molecule to form halogenur mol effect or kharasch effect. hcl and hi do not show the peroxide effect. physical properties: halogenur alkyl being polar in nature are insoluble in water as they cannot break existing H-bonding in water. have higher melting and boiling points. The following is the following order: CH3I > ch3ch2ch2ch2i. the binding force of the bond c x follows the orderch3f > ch3cl > ch3br > CH3li.e., the binding force of the bond c x decreases as the size of the halogen atoms increases. the correct order of rx stability is as follows: > > r br > r i chemical properties dehydrohalogenation : >> Reaction with mg metal : >> Reaction of wurtz: >> Reaction of wurtz: >> r br > r i chemical properties dehydrohalogenation : >> r br > r i chemical properties dehydrohalogenat alchil halides: haloarenes Attached: the general formula is arx, where ar = group aryl, general methods of preparation: for the process of scraping: from the salt of benzene insoluble in water but easily measurable with organic solvents. most of them are steam birds, heavierTheir boiling points are higher than the corresponding alkyl halides. The boiling points gradually rise from fluorine compounds to iod. I give. Property: Â «Nucleophile replacement reactions: the presence of an electron collection group (NO2) in vegetables - and para-positions increases the reactivity of the aloarenes. Electrophyla replacement reactions: Â «Halogenation: A« Nitration: «sulphonation: A« Friedal-Crafts: reaction with metals: suitable reaction: Mechanism. Unimolecular (SN1) Bimolecular (SN2) is a first order reaction. It is a second-order reaction. Generally in protonic polar solvents such as water, alcohol and acetic acid. Executed in polar aprotic solvents such as acetone, dmso, acetonitrile, or dmf.si plays in two phases through carocation as the takes place in a transition phase. The fastest carbocation stability will be the reaction. Rate of reaction. CH3> 1Ã, °> 2Ã, °> 3Ã,° halogen (faster) (higher) minus the stericoin ts obstacle, lâ € \ t ™ action will be faster. Does it proceed with weak nucleophiles, for example CH3O», CN», Ohà ¢ ', etc. The configuration is maintained but In reverse front attack, the inversion of the configuration occurs (reversal of Walden). Å ¢ â ¢ â ¢ â vallyl and benzyl alley halides show greater reactivity in SN1 reactions compared to other primary alkyl halides thanks to the greater stabilization of allyl and benzyl carbonizing intermediates per resonance. to nucleophile replacement reactions. This is due to the double bond character of the CX bond due to resonance. Optical / Enantiomerismoism: Ã ¢ â £ ceThe optical isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the polarized flat light to the left (counterclockwise) and laevoro isomer (Latin: right means right) (D-form or + ve) if it rotates the latin right means right) (D-form or + ve) if it rotates the latin right means right) (D-form or + ve) if it rotates the latin right means right) (D-form or + ve) if it rotates the latin right means right) (D-form or + ve) if it rotates the latin right means right) (D-form or + ve) if it rotates the lati ¢ â ¢ »An equimular mixture of the shape of the form L will optically inactive and is called racemic mixture (or form DL or (Ã, ±) -Meascele) .à ¢ ¢ » The process of converting an enanziomer in a racemic mixture is said to have chirality if the central carbon atom is attached to four different groups and this center is called chiral center (asymmetrical) or stereogenic center o stereocentro. 4 ¢ â ¢ Â » achirality: the mixture is said to have achirality if the coal atom or central has a symmetry plane is aerial (not chiral) and if the molecule does not have a symmetry plane \tilde{A} Chiral. \tilde{A} \hat{c} \hat{A} Enantiomers that are non-superimponable specular images (or asymmetrical) isomers that are non-superimponable specular images (or asymmetry plane \tilde{A} is \tilde{A} These compounds have two or more still number of carbon atoms and have an internal symmetry plan. They are optically inactive due to internal compounds: Composites (1) Cloroformium (CHCl3) Uses - Its main use is in the production of Freon refrigerant, R-22.- It is used as a solvent for resins, tires, oils and fats, alkaloids, iodine and many other substances. In the past, it has been widely used as a result of liver damage. It is used to preserve anatomical species. Effects - It is oxidized to poisonous gas, carbonylchloride, known as phosgen. Phosgen gas causes liver and renidama. - Inhalation of chlorine-shaped vapours depresses the CNS, causes dizziness, fatigue and headaches. (2) Hydeforms (CHI3) Usa - It is used as an antiseptic in clothing of wounds due to the release of iodine. - It is used as a methylation agent in organic synthesis. Effects - It has a strong smell. Freons - They are used as refrigerants, blowing agents, propellants in medical applications and degreesant solvents. - The brakes cause the break of the ozone layer by starting radical reactions of the chain in the stratosphere.- This anthropogenic compound is a greenhouse gas and its effect is more than CO2. (3) DDT Uses - In 1940, it was used as a pesticide. Effects - This is a persistent organic pollutant, strongly absorbed by the soil.- It is lipophilic therefore has a high potential to bioaccumulate. - It can be directly genotoxic but can also induce enzymes to produce other genotoxic intermediates and DNA adducts. The replacement of hydrogen atoms in hydrocarbons, aliphatics or aromatics, from halogen atoms (s) causes the formation of alkyl halide (haloarene,) respectively. Classification of alogen derivatives are classified as mono, tri, tetra, etc. halogen derivatives, for example, On the basis of the carbon nature to which the halogen atom is attached, halogen derivatives are classified as 1°, 2°, 3°, allylic, benzylic, vinyl and arilic derivatives, for example, From Alcohols In the Groove method, ZnC12 is not required. The reactivity order of halogen acids is HI > HBr > HCl. The Darzen procedure is the best method to prepare alchil halides from alcohols since both products (SO2 and HCl) are gaseous and run easily. 2. Alkanes Free Radical Alogenation Addition of Halides Hydrogen on Alkenes 1. Finkelstein Action 2. Reaction Swarts H3C - Br + AgF + \hat{a} \hat{a} CR3BR> CH3I 3. The time of the dipole decreases as the electronegativeness of alogen decreases. Core replacement reactions (SN reactions) KCN is mainly Ionic and provides cyanide ions in solution, which is ambiguous nucleophile and bond with the carbon side to form as the main product, while AGNO2 produces R-NO2 as a product. Vinyl chloride is less responsive to nucleophile replacement reactions due to resonance. The nucleophile replacement reactions due to resonance are fixed in two steps: rate, r = k [rx]. It is a reaction of the first order. Order of reactivity of the alkyl object towards the mechanism SN1 3 °> 2 °> 1 ° Polar polar solvents, low nucleophile concentration and weak nucleophiles Book the mechanism SN1. In the SN1 reactions, the partial race occurs due to the possibility of frontal attack and back on the planar carbocsation. (b) Type SN2 (bimolecular nucleophile replacement) These reactions proceed in one step and is a reaction of the second order with ... K [RX] [NU]. During the SN2 reaction, the reversal of the configuration occurs (reversal of walden) i.e., starting from a dextrorotory halide is obtained a Laevo product and vice versa, for example, the reactivity of alogenurs towards the SN2 mechanism is 1 °> 2 °> 3 ° TASSO DI REATION IN SN2 The mechanism depends on the force of the attacking nucleophile. The strength of some common nucleophiles is: cn㠀 ">: i⠀">: o ...>: Oh†"> concentration Book the SN2 mechanism. The relative reactivity of some alkyl in the SN1 and SN2 reactions are in order, the resonance structure of benzil carboocations is relative reactivity of some alkyl in the SN1 and SN2 reactions are in order, the resonance structure of benzil carboocations is relative reactivity of alkyl alkyl for the same alkyl group is re> rbr> rci> rf 2. Dehydrohalogenation between halogenurs 3 °> 2 °> 1 ° 3. Reduction 4. Reaction with metals with Agent Grignard is never isolated in the solid state as it explodes in the dry state. So it is used as an ethereal solution. 5. General methods of preparation of aryl halides 1. For halogenation of aryl halides 1. For halides 1. For halogenation of aryl halides 1. For ha Diazonium 4. From the physical properties of Phenol of the aryl halides 1. The aryl halides are colorless or of the P-ISOMER A It is more than 0- and M-Isomer. This is due to the most symmetrical nature of the P-Isomer. 4. Due to the resonance in chlorobenzene, the link C-ci is brief and therefore, its momento is lower than that delcyclohexylelloride. Chemical properties of aryl halides 1. ARYLIDES replacement nucleophile replacement reactions are less reactive towards the nucleophilic replacement reaction. Their low reactivity is attributed due to the following reasons: due to the molecule by delocalizing electrons. (Instability of phenyllic carbulas. However, aryl halides having electron collection groups (such as à ¢ â, "No2, -So3h, etc.) to vegetable garden positions and para does not easily suffer a nucleophile replacement reaction. The presence of Electrophilic replacement reaction of the craftsmanship Friedel confer to one Mixture of o- and p- chlorine derivatives â €

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